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| Al has a higher charge (+3) than sodium (+1)  More delocalised electrons  Stronger metallic bonding |
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| Why does Aluminium not follow the general trend in first ionisatipn energies across the period? |
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| Why would an element be classed as a p-block element? |
| The outer electron is in a p orbital |
|  |
| Why would an element be classed as a s-block element? |
| The outer electron is in a s orbital |
|  |
| Why would an element be classed as a d-block element? |
| The outer electron is in a d orbital |