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| State why the Maxwell Boltzmann curve goes through (0,0) |
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| With reference to the Maxwell-Boltzmann distribution, explain why a decrease in temperature decreases the rate of reaction |
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| None |
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| Collision between the reactant particlesWith the activation energy |
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| A small increase in temperature results in many more molecules having energy greater than the activation energy |
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| The minimum energy for a reaction to occur |
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| Explain the process that causes some molecules to have very low energies |
| Collisions cause some molecules to slow down/lose energy |
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| Suggest a reason why a solid catalyst is often in the form of a powder. |
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| In terms of the behaviour of particles, explain why the rate of reaction decreases with time |
| Fewer collisions as there are less reactant particles |
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| What does the area under s Maxwell-Boltzmann curve represent? |
| The total number of particles |