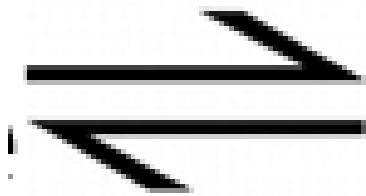


<p>What is a reversible reaction?</p>	<p>What arrow is used to represent reversible reactions?</p>
<p>If the forward reaction is exothermic, what will the backwards reaction be?</p>	<p>What is meant by equilibrium?</p>
<p>At equilibrium, what is special about the concentrations of the reactants and products?</p>	<p>What is meant by the position of equilibrium?</p>
<p>Equilibrium only occurs in a closed system. What is meant by a closed system?</p>	<p>What is meant by the position of equilibrium being on the left?</p>



Where the products can be changed back into the reactants.

Rate of the forward reaction = the rate of the backward reaction

Endothermic

The proportion of products and reactants at equilibrium

They remain the same

There are more reactants than products

Where no chemicals can enter or leave e.g. a stoppered flask

<p>How does using a catalyst affect the position of equilibrium?</p>	<p>What is Le Chatelier's principle?</p>
<p>How does increasing the temperature of a reaction affect the position of equilibrium?</p>	<p>How does increasing the amount or concentration of reactants affect the position of equilibrium?</p>
<p>How does increasing the pressure affect the position of equilibrium?</p>	<p>What is meant by the equilibrium yield?</p>
<p>Why is a compromise temperature sometimes used?</p>	<p>Why is a compromise pressure sometimes used?</p>

When a change is made to a reaction at equilibrium, the position of equilibrium moves to oppose the change	It doesn't
It favours the products side of the reaction	It favours the endothermic reaction
How much of a product is made at equilibrium	It favours the side with the fewest moles of gas
High pressure is expensive, uses a lot of energy and needs specialist equipment	A low temperature will give a high yield if the forward reaction is exothermic, but it will also give a slow rate of reaction