Draw the symbol that represents a reversible reaction in a chemical equation

In dynamic equilibrium, what does the dynamic part mean?

C5 Monitoring and Controlling Chemical Reactions

C5.2 Equilibria

Links

← 3.1 Introduction to Chemical Reactions
5.1 Controlling Reactions

→ 6.1 Improving Processes and Products

Define equilibrium

What happens to the position of equilibrium if the temperature is increased?

What happens to the position of equilibrium if the pressure is increased?

What is the exception to this and what happens instead?

What do we mean by the position of equilibrium?

What is a reversible reaction?

How is the position of equilibrium affected by the addition of a catalyst?

How will removing the products from the reaction vessel affect the position of equilibrium?

How will increasing the concentration of reactants affect the position of equilibrium?

Why is the pressure we use sometimes lower than the ideal pressure?

What is a compromise condition?

Why is the temperature we use sometimes higher than the ideal temperature?

Keywords