

<p>What is the limiting reactant?</p>	<p>What do we mean when we say a reactant is in excess?</p>
<p>What is the yield of a reaction?</p>	<p>What is the theoretical yield of a reaction?</p>
<p>What is the formula for calculating percentage yield?</p>	<p>Why is a high percentage yield important?</p>
<p>What is the formula for calculating atom economy?</p>	<p>Why is a high atom economy important?</p>

<p>The reactant that is left over after the reaction has finished</p>	<p>The reactant that gets used up first</p>
<p>The maximum amount of product made if all the reactants atoms were converted to products</p>	<p>The mass of a product made in a reaction</p>
<p>Reduces costs Doesn't waste starting material</p>	<p><math>(\text{actual yield} / \text{theoretical yield}) \times 100</math></p>
<p>Reduces unwanted products, more sustainable and maximises profit</p>	<p><math>(\text{sum of the Mr of the desired product} / \text{sum of the Mr of all products}) \times 100</math></p>

<p>What is the formula for calculating concentration in mol/dm<sup>3</sup>?</p>	<p>What factors need to be considered when choosing a reaction pathway?</p>
<p>What is the formula for calculating concentration in g/dm<sup>3</sup>?</p>	<p>How do you convert cm<sup>3</sup> to dm<sup>3</sup>?</p>
<p>How do you convert concentration of a solution from mol/dm<sup>3</sup> to g/dm<sup>3</sup>?</p>	<p>What is the end-point of a titration?</p>
<p>Why is universal indicator not used in titrations?</p>	<p>How much volume does one mole of gas occupy at room temperature and pressure?</p>

Percentage yield, atom economy, whether the by-product is useful or difficult to get rid of, rate of reaction, equilibrium position	Amount of solute in moles / volume of solution in $\text{dm}^3$
Divide by 1000	Amount of solute in grams / volume of solution in $\text{dm}^3$
When the indicator changes colour	Multiply the concentration by the Mr
$24\text{dm}^3$	It gives a gradual colour change whereas a single indicator gives a sudden colour change