What is the test for oxygen?	What is the test for hydrogen?
What is the test for chlorine?	What is the test for carbon dioxide?
List the steps involved in carrying out a flame test for metal ions?	What colours do lithium, sodium, potassium, calcium and copper give in a flame test?
What colour precipitate forms when iron (II), iron (III), copper, calcium and zinc solutions are added to sodium hydroxide?	What is the test for sulfate ions?

Gives a squeaky pop with a lit splint	Relights a glowing splint
Turns limewater from clear to cloudy	Turns damp blue litmus paper red then white
Lithium = red Sodium = yellow Potassium = lilac Calcium = orange/red Copper = green/blue	1. Clean a nichrome wire loop with HCl and a blue Bunsen flame 2. Dip the loop into the sample 3. Hold the loop in a blue Bunsen flame 4. Record the colour of the flame
Add HCl, then barium chloride. If positive a white precipitate will form	Iron (II) - green Iron (III) - orange/brown Copper—blue Calcium—white Zinc—white, redissolves in excess

What is the test for carbonate ions?	What is the test for halide ions?
What colour precipitate do chloride, bromide and iodide ions give with silver nitrate?	What advantages do scientific instruments have when performing tests for chemicals?
What does each peak on a gas chromato-gram represent?	What does the areas under a peak in a gas chromatogram repre-sent?
What information does a mass spectra give?	For the mass spectra of a compound which peak gives us the mass of the molecule?

Add nitric acid then silver nitrate, record the colour of the precipitate	Add hydrochloric acid. If positive bubbles of carbon dioxide will be given off.
Sensitive, accurate, fast and can run continuously	Chloride—white Bromide—cream Iodide—yellow
The amount of the substance present	A different chemical
The peak furthest to the right	The mass of an atom or molecule