

<p>What type of compound is used as the electrolyte in electrolysis?</p>	<p>What does PANIC stand for in electrolysis?</p>
<p>What two properties must the material chosen for the electrode have?</p>	<p>What happens to positive ions in electrolysis?</p>
<p>What happens to negative ions in electrolysis?</p>	<p>Which electrode are metals (and hydrogen) attracted to?</p>
<p>Which electrode are non-metals (except hydrogen) attracted to?</p>	<p>Why can't solid ionic compounds be used for electrolysis?</p>

Positive Anode Negative Is Cathode	A molten or dissolved ionic compound
They are attracted to the cathode where they gain electrons	Unreactive and a good conductor of electricity
Cathode	They are attracted to the anode where they lose electrons
Ions cannot move	Anode

<p>In electrolysis of solutions, what is discharged at the cathode if the metal is less reactive than hydrogen?</p>	<p>In electrolysis of solutions, what is discharged at the cathode if the metal is more reactive than hydrogen?</p>
<p>In electroplating, which electrode becomes coated in metal?</p>	<p>In electroplating, which electrode is made from the metal that you want to coat the object in?</p>
<p>In electroplating, what must the electrolyte contain?</p>	<p>What will happen to the mass of the cathode during electroplating?</p>
<p>What will happen to the mass of the anode during electroplating?</p>	<p>What is discharged at the anode in the electrolysis of solutions?</p>

Hydrogen	The metal
Anode	Cathode
It will increase	Ions of the metal that you are coating the object with
A halogen if the anion is a halide ion, otherwise oxygen	It will decrease