

What is oxidation?	What is reduction?
What is an oxidising agent?	What is a reducing agent?
What ions are generated by acids?	What ions are generated by alkalis?
Which is the best method of measuring pH and why?	What two substances are made when an acid reacts with an alkali?

The gain of electrons	The loss of electrons
A chemical that causes another chemical to gain electrons	A chemical that causes another chemical to lose electrons
OH^-	H^+
A salt + water	pH meter as universal indicator gives a gradual colour change

<p>What 3 substances are made when a metal carbonate reacts with an acid?</p>	<p>What is the ionic equation for the reaction of an acid with an alkali?</p>
<p>What two substances are made when an acid reacts with a metal?</p>	<p>What is the difference between a dilute acid and a concentrated acid?</p>
<p>What is the difference between a weak and a strong acid?</p>	<p>What does a high pH mean in terms of concentration of H^+ ions?</p>
<p>What does a low pH mean in terms of concentration of H^+ ions?</p>	<p>A decrease of 1 in the pH scale means that the concentration of H^+ ions has increased by how much?</p>

$\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$	A salt, water and carbon dioxide
A dilute acid has few molecules of acid per unit volume, a concentrated acid has many molecules of acid per unit volume	A salt + hydrogen
A low concentration of H^+ ions	A strong acid completely splits up into its ions in water, a weak acid only partially splits up into its ions in water
10x	A high concentration of H^+ ions