

What does diatomic mean?	Which elements are diatomic?
What is the formula for the ammonium ion?	What is the formula of a hydroxide ion?
What is the formula of a nitrate ion?	What is the formula of a carbonate ion?
What is the formula of a sulfate ion?	What are the tests for hydrogen, oxygen, carbon dioxide and chlorine?

<p>Nitrogen, oxygen, hydrogen, bromine, chlorine, iodine, fluorine</p>	<p>A molecule which contains two atoms joined together</p>
<p>OH^-</p>	<p>NH_4^+</p>
<p>CO_3^{2-}</p>	<p>NO_3^-</p>
<p>Hydrogen—Squeaky pop with lit splint Oxygen—relights a glowing splint Chlorine—turns damp blue litmus paper red then white Carbon dioxide—turns limewater from clear to cloudy</p>	<p>SO_4^{2-}</p>

<p>What are the state symbols for solid, liquid, gas and aqueous?</p>	<p>What is a spectator ion?</p>
<p>What is a mole?</p>	<p>How much does a mole of an element weigh?</p>
<p>How much does a mole of compound weigh?</p>	<p>What is a limiting reactant?</p>
<p>What is meant by a reactant that is in excess?</p>	<p>What is the formula for calculating concentration in mol/dm³ and g/dm³?</p>

<p>An ion that is identical on both sides of a chemical equation</p>	<p>Solid = (s) Liquid—(l) Gas = (g) Aqueous = (aq)</p>
<p>The relative atomic mass in grams</p>	<p>A mole is 6.022×10^{23} particles of a substance</p>
<p>The reactant that gets used up first, causing the reaction to stop</p>	<p>The relative formula mass in grams</p>
<p>Amount of solute in moles per volume of solution in dm^3 / Amount of solute in grams / volume of solution in dm^3</p>	<p>The reactant that is left over at the end once the reaction has stopped</p>