

<p>What is an allotrope?</p>	<p>Why is diamond hard and has a high melting point?</p>
<p>Why does graphite conduct electricity?</p>	<p>Why does diamond not conduct electricity?</p>
<p>Why is graphite slippery?</p>	<p>What are the potential uses of fullerenes?</p>
<p>Why does graphite have a high melting point?</p>	<p>Why are metals malleable?</p>

<p>It has many strong covalent bonds</p>	<p>Different forms of the same element (different arrangement of atoms)</p>
<p>It has no delocalised electrons</p>	<p>It has delocalised electrons</p>
<p>Drug delivery around the body, an industrial lubricant</p>	<p>Weak forces between the layers so they can slide over each other</p>
<p>The layers of positive ions can slide over each other</p>	<p>Many strong covalent bonds need to be broken</p>

<p>What name is given to a solid turning into a liquid?</p>	<p>What name is given to a liquid turning into a solid?</p>
<p>What name is given to a gas turning into a liquid?</p>	<p>What name is given to a liquid turning into a gas?</p>
<p>Why do metals conduct electricity?</p>	<p>What is the name given to a solid turning into a gas?</p>
<p>When do ionic compounds conduct electricity?</p>	<p>How many covalent bonds does carbon normally form?</p>

Freezing	melting
evaporation	condensation
sublimation	They have delocalised electrons
4	When molten (melted) or dissolved