

<p>What are the relative masses and charges of electrons, neutrons and protons?</p>	<p>What is the atomic number of an element?</p>
<p>Which subatomic particles are found in the nucleus?</p>	<p>Define an isotope</p>
<p>Define the mass number</p>	<p>Define relative atomic mass</p>
<p>Which subatomic particle is found in shells orbiting the nucleus?</p>	<p>How do we work out how many electrons an atom has?</p>

<p>Number of protons in one atom or nucleus</p>	<p>Electron—mass 1/2000, charge -1 Neutron—mass 1, charge 0 Proton—mass 1, charge +1</p>
<p>Atoms with the same number of protons and different numbers of neutrons</p>	<p>Protons and neutrons</p>
<p>Average mass of 1 atom of an element compared to 1/12 mass of one atom of ^{12}C</p>	<p>The number of protons and neutrons</p>
<p>Equal to the atomic number</p>	<p>Electron</p>

<p>How do we work out how many protons an atom has?</p>	<p>How do we work out how many neutrons an atom has?</p>
<p>How do we work out how many electrons a positive ion has?</p>	<p>How do we work out how many electrons a negative ion has?</p>
<p>What did John Dalton contribute to atomic theory?</p>	<p>What did JJ Thompson discover?</p>
<p>What did the gold foil experiment by Rutherford, Geiger and Marsden lead to?</p>	<p>What did Bohr contribute to atomic theory?</p>

Mass number minus the atomic number	Equal to the atomic number
Atomic number plus the charge on the ion	Atomic number minus the charge on the ion
The electron	Atoms are indivisible, all atoms of one element are identical.
Electrons occupied fixed shells	The idea of a nucleus in an atom